

1. T, T, T, T, F, F, T, F, F, T
2. dipole-dipole, dispersion (strongest)
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3. 2.15
4. 32.0%
5. covalent network
metallic network
6. -81.35
7. 4.97
8. hydrogen bonding, dipole-dipole, dispersion
dipole-induced dipole, dispersion
9. 0.364
10. weaker, CH₄ stronger, C₂H₆
Both have dipole-induced dipole and dispersion. C₂H₆ has stronger IFs with water because it has more atoms (OR more surface area).
less opposed by S, CH₄ more opposed by S, C₂H₆
C₂H₆ is larger and {more hydrophobic OR is more shunned by water OR interferes more with hydrogen bonding}.
11. Cl₃PO
H₂O
12. 0.356